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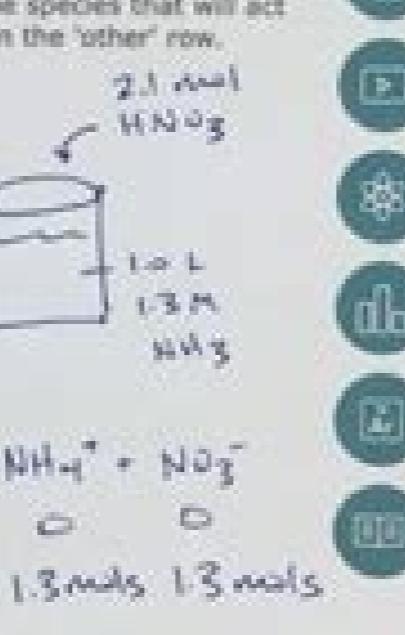
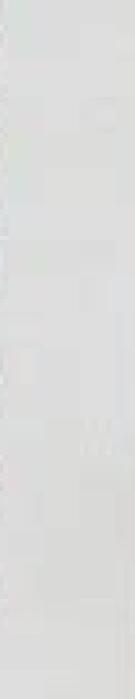
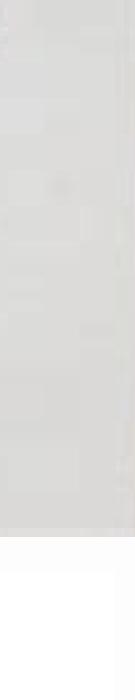
6 Strong Acids		6 Strong Bases	
HClO <sub>4</sub>	perchloric acid	LiOH	lithium hydroxide
HCl	hydrochloric acid	NaOH	sodium hydroxide
HBr	hydrobromic acid	KOH	potassium hydroxide
HI	hydroiodic acid	Ca(OH) <sub>2</sub>	calcium hydroxide
HNO <sub>3</sub>	nitric acid	Sr(OH) <sub>2</sub>	strontium hydroxide
H <sub>2</sub> SO <sub>4</sub>	sulfuric acid	Ba(OH) <sub>2</sub>	barium hydroxide

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sulfuric acid	strontium hydroxide
	Ba(OH) <sub>2</sub>
	barium hydroxide

The preparations of two aqueous solutions are described in the table below. For each solution, write the chemical formulas of the major species present at equilibrium. You can leave out water itself.

Write the chemical formulas of the species that will act as acids in the 'acids' row, the formulas of the species that will act as bases in the 'bases' row, and the formulas of the species that will act as neither acids nor bases in the 'other' row.

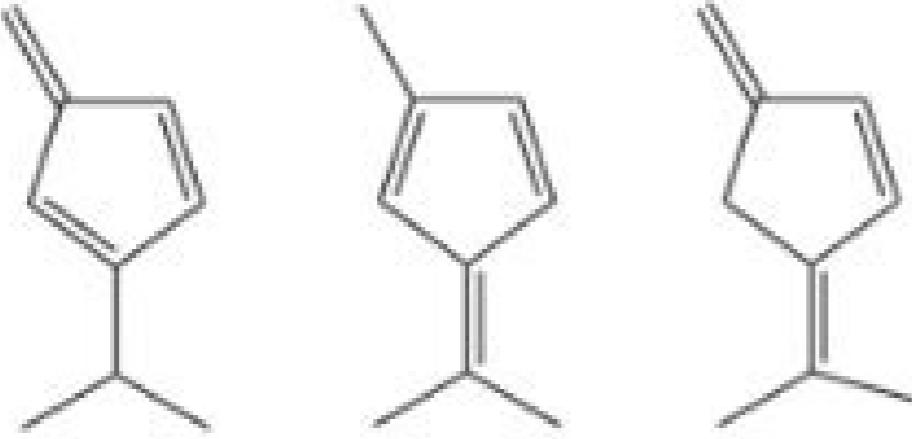
You will find it useful to keep in mind that NH<sub>3</sub> is a weak base.

2.1 mol of HNO <sub>3</sub> is added to 1.0 L of a 1.3 M NH <sub>3</sub> solution.	X acids: <input type="checkbox"/> HNO <sub>3</sub> , NH <sub>4</sub> <sup>+</sup> <input type="checkbox"/> bases: <input checked="" type="checkbox"/> X other: <input type="checkbox"/> NO <sub>3</sub> <sup>-</sup>		
0.1 mol of KOH is added to 1.0 L of a solution that is 0.6 M in both NH <sub>3</sub> and NH <sub>4</sub> Br.	<input type="checkbox"/> acids: <input checked="" type="checkbox"/> <input type="checkbox"/> bases: <input type="checkbox"/> <input type="checkbox"/> other: <input type="checkbox"/>	$\text{NH}_3 + \text{HNO}_3 \rightarrow \text{NH}_4^+ + \text{NO}_3^-$ 1.3 mol 2.1 mol O O exch: O 0.8 mol 1.3 mol 1.3 mol	

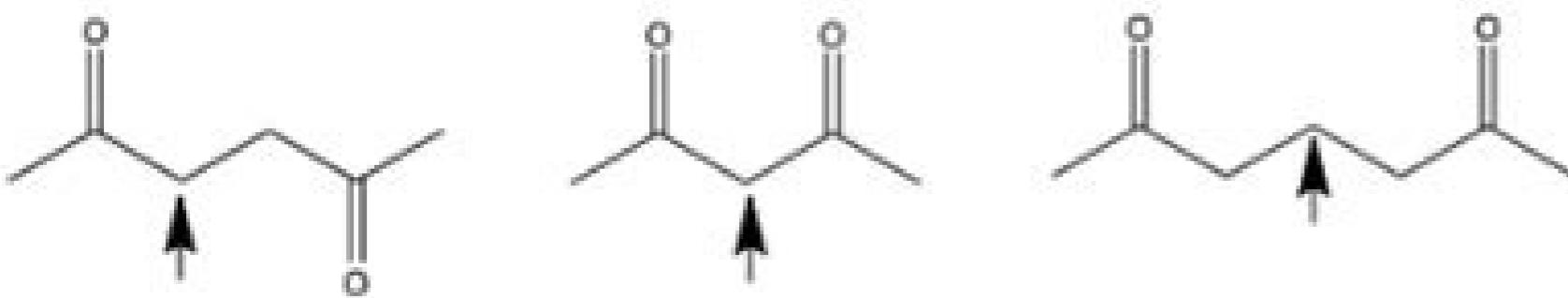
$\text{NH}_3 + \text{KOH} \rightarrow \text{NH}_4^+ + \text{K}^+$

$\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{NH}_3 + \text{H}_3\text{O}^+$

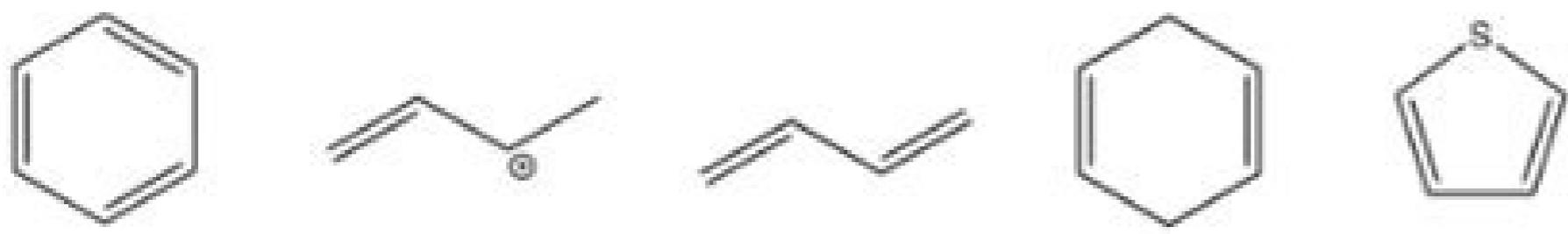
### 5. Which of the following is the strongest acid and why?



### 6. Rank the indicated H's from most to least acidic.



### 7. Which of the following have delocalized electrons?



Compare the conjugate bases of these three acids.

Acid 1: sulfuric acid , H<sub>2</sub>SO<sub>4</sub>

Acid 2: hypochlorous acid , HClO

Acid 3: hydrogen sulfide , HS<sup>-</sup>

What is the formula for the weakest conjugate base ?

Submit Answer

Retry Entire Group

9 more group attempts remaining

H2BO3- ion dihydrogen carbonate  $4.7 \times 10^{-11}$  hydrogen carbonate ion HCO3- Carbonate ion 4.2 \* 10-13 hydrogen phosphate HPO4 2- ion phosphate 1.8 \* 10-13 ion dihydrogen borate H2BO3- HBo3 2- ion hydrogen borate 1.3 \* 10-13 ion hydrogen sulfide HS-s 2-sulfuro 1.6 \* 10-14 Hydrogen ion HBO3 2- bo3 3- ion borato ----- water H2O Basicity of basicity H3PO4 is defined as the number of hydromassage ogen atoms present in an acid. The conjugated base of H3PO4 is H2PO4- and H3PO3 is H2PO3- and according to the resonance structures, the former conjugated base is more stable. He's a poor rusty agent. Phosphoric acid or orthophosphoric acid is a tricotic acid that contains phosphorous with the chemical formula H3PO4. So, according to Arrhenius' theory, he acts like an acid. 3. H3PO3 is a stronger acid than H3PO4 due to experimentally calculated pKa values: the pKa value of H3PO3 (1.3) is lower than H3PO4 (2.1) which indicates that the first is a stronger acid. afe It is used as a disinfectant agent, dispersion agent, electrolyte, etc. By the way, you should also read the article on the structure of Lewis H3PO4. However, the pKa value of H3PO4 is about 2.14, which again indicates it is a day with a lower capacity to donate a proton in an aqueous solution. Is H3PO4 a strong or strong? Demne erito: The compound compounds do not explain how BF3 and AlCl3 show acid properties without hydrogen. We understand the nature of H3PO4 (acidic or basic) with the help of previous theories. Theory of Arrhenius for acid: When H3PO4 dissolves in water, it breaks down into two constituent ions, H2PO4- and H+. In this article we will discuss some of the most sought-after questions related to phosphorus-like acid • Nature of force (force or weak) • H3PO4 conjugated couple • Acidity or Basicity of H3PO4 So, is H3PO4 an acid or base? Thank you so much for your cooperation. It contains four oxygen atoms, a phosphorus atom and three hydrogen atoms. The strong bases are listed at the bottom right of the table and are further weakened as we move to the top of the table. Compound name Fosphorus acid Chemical formula H3PO4 Structure 2D - 3D • Nature Acidic (mode) pKa 2.14 Couple conjugated H2PO4- Why does H3PO4 act like an oxygen acid? Therefore, it remains undissociated in an aqueous solution and is considered a weak acid. The wrong went wrong. The lower the pKa value, the higher the acid character of that compound. H3PO4 + H2O → entitled H3O+ + H2PO4- is it phosphoric (H3PO4) a strong acid or weak acid? Third step: In this step, H2PO4- donates the second acid proton to the water to form PO43-. Two main theories in the chemistry that define acids and bases are: • Theory of conjugation theory Bronsted-Lowry Defines an acid as a donor of H+ in a solution. Defines an acid as a donor of H+ in a solution. Defines an acid as a donor of protons to form the conjugated base Merit: Explains the reaction between acids and bases that produce water and salts. A strong acid always forms a conjugated base. Donate its hydrogen Ejido in an aqueous solution to the elision of its conjugated base H2PO4- è è Conjugated base stability: The most stable is the conjugated base of an Ejido, the greatest one will be the fade. Similarly, other possible conjugated bases such as HPO42- è è and PO43- è è They also react with hydrogen ions to make PO43- remain in the acidity order determined by each subsequent dissociation. In the case of the Ejido conjugate, there is no difference. In the case of H3PO4, for the H2PO4- (dihydrogen phosphate) conjugated base after the loss of a dihydrogen proton. When the hydrogen binds to another electron pair such as three oxygen then it is considered Ejido. The statement that the strongest ion is completely in an aqueous solution is one of the unqualification of bil ioniza or dissociate partially in a solution. A base of Bil always forms a strong spouse. Again, H2PO4- It is a day than Bil that H2PO4- and H3PO4 (indicating a higher pKa value of 12.37) so it is made more diffile for H2PO4- Eliminate the third proton © ethical in an aqueous solution. So, according to Bronsted-Lowry, it is an Ejido. The H3PO4 dissociation in aqueous solution is as follows: H3PO4 + H2O → H3O+ + H2PO4- [PKA = 2.14] H2PO4- + H2O → H3O+ + HPO42- [PKA = 7.20] HPO42- + H2O → H3O+ + PO43- [PKA = 12.37] First step: In this step, H3PO4 DON Water to form H2PO4- but H2PO4- is a more reducing agent. The commercial method of preparation of phosphoric acid is the wet process in which sulfuric acid gets added to phosphate rock in a chain of well-stirred reactors. H3PO4 is a weak acid because it does not dissociate completely in its aqueous solution or water. The basicity of H3PO4 is three (3). It appears as a colorless, odorless syrupy liquid or transparent crystalline solid. In H3PO4, there are three O-H bonds or three acidic hydrogen atoms hence, it is a tribasic acid. Wait a moment and try again. It is used in compound semiconductor processing. Why is H3PO4 a weaker acid than H3PO3? It is used as an additive and flavoring agent in animal or poultry feed. It is used as teeth whiteners or mouth washing liquids. A weak acid always forms a strong conjugate base. Reducing property The oxidation state of Phosphorus in H3PO4 is +5 which cannot be oxidized further. So, H3PO4 is a weaker acid than H3PO3. Polarity: H3PO3 is more polar than H3PO4. So, H3PO4 is weaker than H3PO3. It is used as a rust remover from the surfaces of various metals like steel, iron, etc. Moreover, H3PO4 is a weak acid so according to the conjugate acid-base pair, its corresponding pair, H2PO4- is a strong conjugate base. Please share it with your friends in your school. It is present in our bones and teeth and acts as a metabolite. P-OH bond: The more the P-OH bond, the lower the acidic strength of the oxyacid. In order to continue enjoying our site, we ask that you confirm your identity as a human. Hence, H3PO3 is a stronger acid than H3PO4. H3PO4 (aq) → H+ + H2PO4- (aq) Bronsted-Lowry theory for acid: When H3PO4 is dissolved in water, it donates one proton to the H2O to form a hydronium J I'm gonna go :P sa a scitemsoc ni desu si ti ¢ sead .3/4 s4OP2H, esab etagunoc a semoceb flesi dna

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