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Is h3po4 a strong acid or weak acid list pdf

6 Strong Acids		6 Strong Bases	
HClO ₄	perchloric acid	LiOH	lithium hydroxide
HCl	hydrochloric acid	NaOH	sodium hydroxide
HBr	hydrobromic acid	KOH	potassium hydroxide
HI	hydroiodic acid	Ca(OH) ₂	calcium hydroxide
HNO ₃	nitric acid	Sr(OH) ₂	strontium hydroxide
H ₂ SO ₄	sulfuric acid	Ba(OH) ₂	barium hydroxide

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The preparations of two aqueous solutions are described in the table below. For each solution, write the chemical formulas of the **major** species present **at equilibrium**. You can leave out water itself.

Write the chemical formulas of the species that will act as acids in the 'acid' row, the formulas of the species that will act as bases in the 'bases' row, and the formulas of the species that will act as neither acids nor bases in the 'other' row.

You will find it useful to keep in mind that NH_3 is a weak base.

2.1 mol of HNO_3 is added to 1.0 L of a 1.3 M NH_3 solution.	X acids: HNO_3 , NH_4^+ bases: X other: NO_3^-	
0.1 mol of KOH is added to 1.0 L of a solution that is 0.6 M in both NH_3 and NH_4Br .	acids: bases: other:	$\text{NH}_3 + \text{HNO}_3 \rightarrow \text{NH}_4^+ + \text{NO}_3^-$ $1.3 \text{ mols} \quad 2.1 \text{ mols} \quad 0 \quad 0$ encl: 0 0.8 mols 1.3 mols 1.3 mols

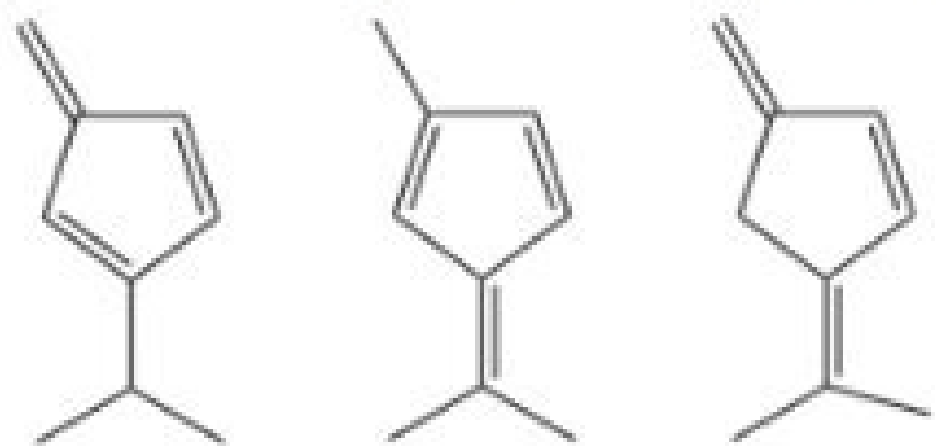
Handwritten notes for the second row:

Diagram of a beaker containing 1.0 L of a solution that is 0.6 M in both NH_3 and NH_4Br . Above the beaker, an arrow points to a label KOH . Inside the beaker, the labels NH_3 and $\text{NH}_4^+ \text{ Br}^-$ are written.

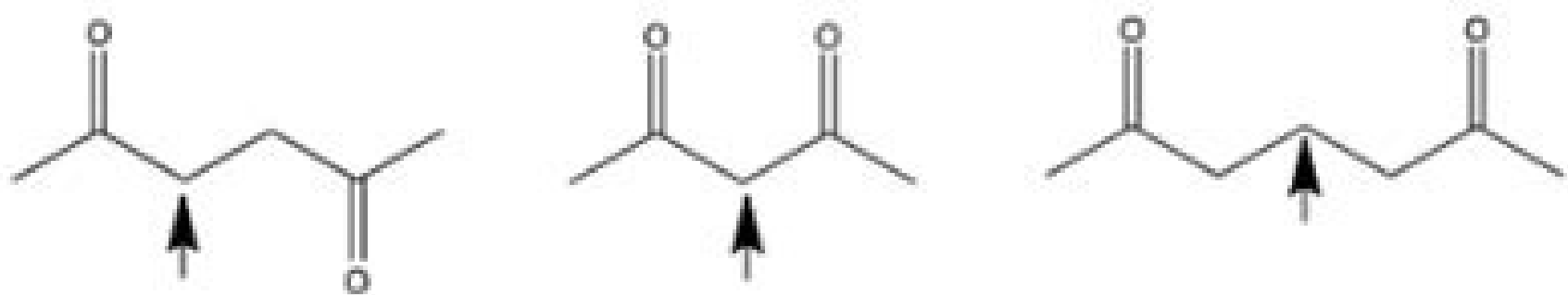
Chemical reaction:

$$\text{KOH} + \text{NH}_4^+ \rightarrow \text{H}_2\text{O} + \text{NH}_3 + \text{K}^+$$

5. Which of the following is the strongest acid and why?



6. Rank the indicated H's from most to least acidic.



7. Which of the following have delocalized electrons?



Compare the **conjugate bases** of these three acids.

Acid 1: **sulfurous acid** , H_2SO_3

Acid 2: **hypochlorous acid** , HClO

Acid 3: **hydrogen sulfide** , HS^-

What is the formula for the **weakest conjugate base** ?

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Is H_3PO_4 a strong acid or weak acid. Is H_3PO_4 strong acid. Is H_3PO_4 strong or weak acid.

[illegible]

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